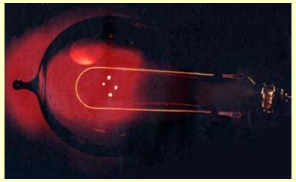
Generation of light – Light sources



Black-body radiation

Luminescence

Laser

Luminescence

Repetition

Types of energy states in atoms and molecules are independent (not coupled)

Energy states are non-continuous, but discrete

Transition between states involves packets (quanta) of energy)

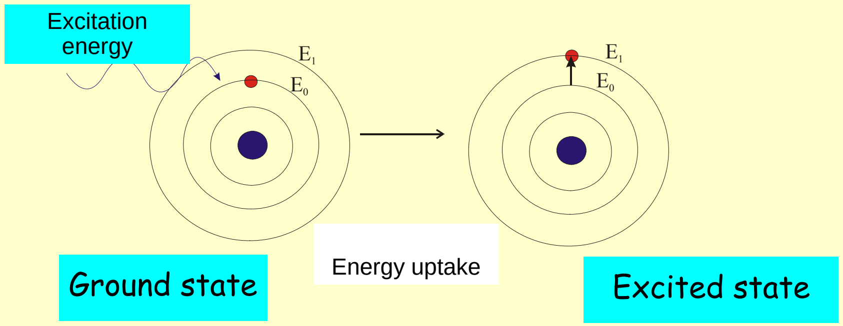
$$E_{total} = E_e + E_v + E_r$$

Scales of transition energies between different states are different:

$$E_e > E_v > E_r$$

Consider a single atom

- Energy states are discrete
- Electrons occupy the lowest possible energy state (ground state)
- Pauli exclusion principle: no two identical fermions (particles with half-integer spin) may occupy the same quantum state simultaneously



Excitation modes

- absorption of radiation (UV/VIS) : *photoluminescence*

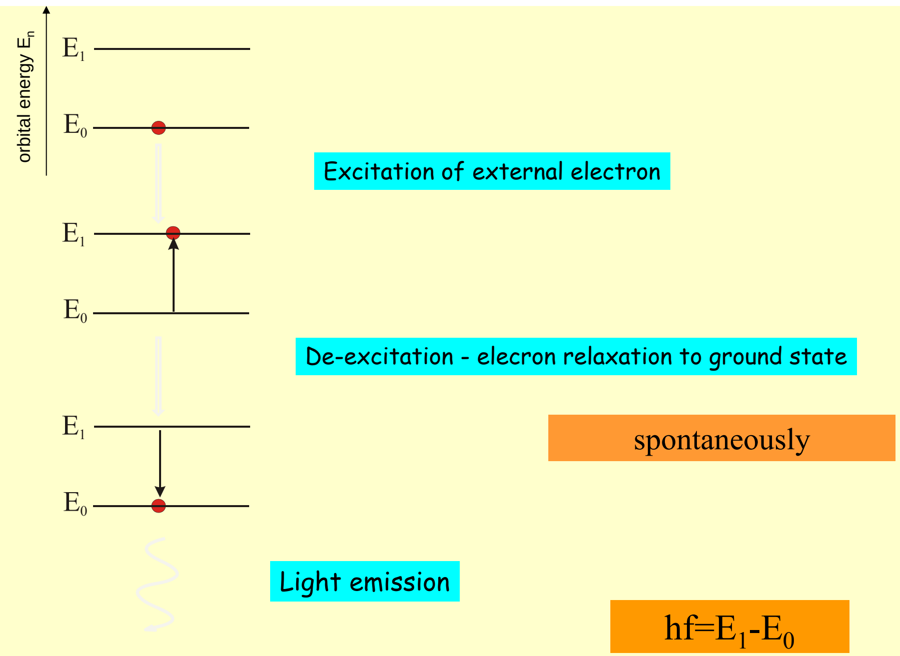
- chemical reaction: *chemo/bio-luminescence*

- Injection of charges: *electroluminescence*

- friction (mechancal deformation: *triboluminescence*

-thermally activated ion recombination: *thermoluminescence*

- Sound waves: *sonouminescence*

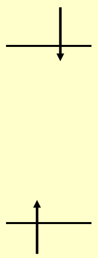


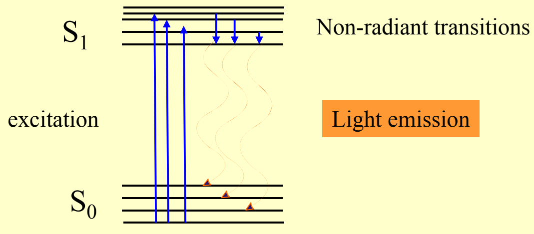
Luminescence: spontaneous light photon emission by electrons when they return from their excited state to their original (ground) state of lower energy

$$hf=E_1-E_0$$

The emitted photon energy is characteristic for the electronic orbitals, thus for the atom/molecule.

The energy of the electronic orbitals in molecules is perturbed by the discrete states of molecular vibrations





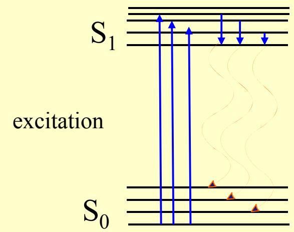
Fluorescence

First excited singlet state

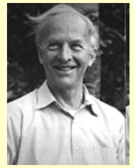
the spin states of the electrons

sum of all e⁻ spin quantum numbers=0

De-excitation by photon emission between singlet states



Kasha's rule:



fluorescence originates always from the vibrational state of lowest energy within the lowest electronic excited state.

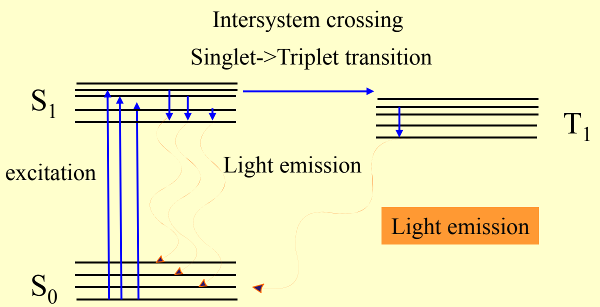
Fluorescence

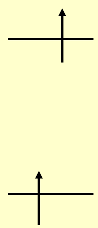
De-excitation by photon emission between singlet states

$E_{excitation} \geq E_{fluorescence}$

$\lambda_{excitation} \leq \lambda_{fluorescence}$

Stokes-shift





Phosphorescence

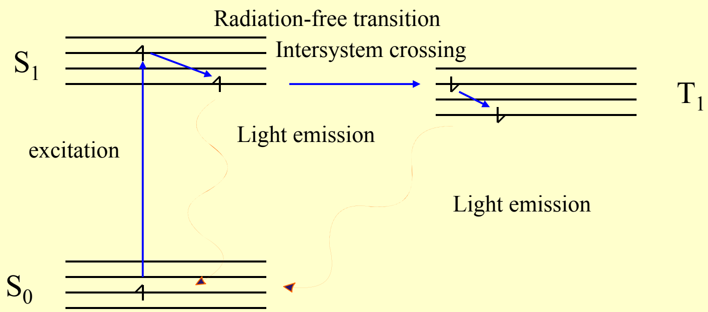
spontaneous light photon emission from metastable excited state

Triplet state

sum of all e⁻ spin quantum numbers=1

Metastable state

Emitted photon energies



Fluorescence **Phosphorescence**

Stokes-shift

$E_{excitation} \geq E_{fluorescence} > E_{phosphorescence}$

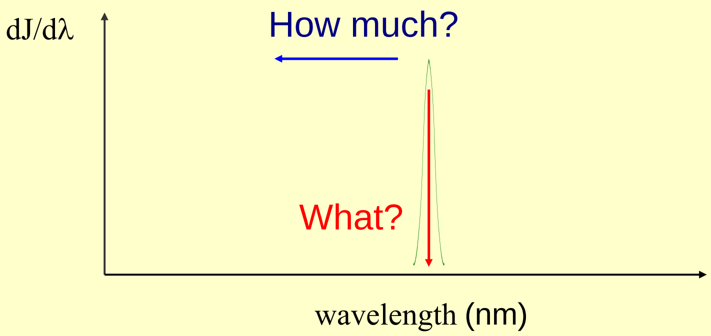
$\lambda_{excitation} \leq \lambda_{fluorescence} < \lambda_{phosphorescence}$

Characteristics of emitted light

Wavelength distribution of emitted light

Emission spectrum

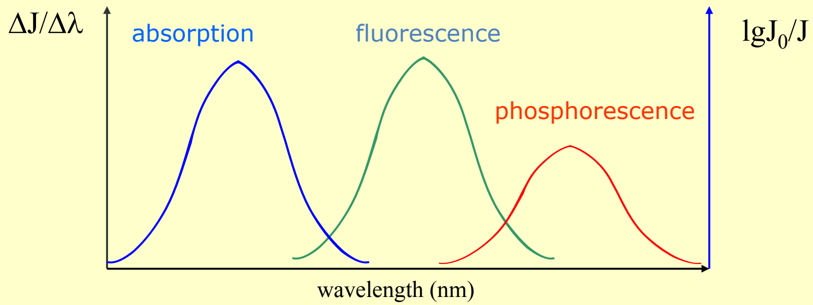
Atoms line spectrum



Wavelength distribution of emitted light

Emission spectrum

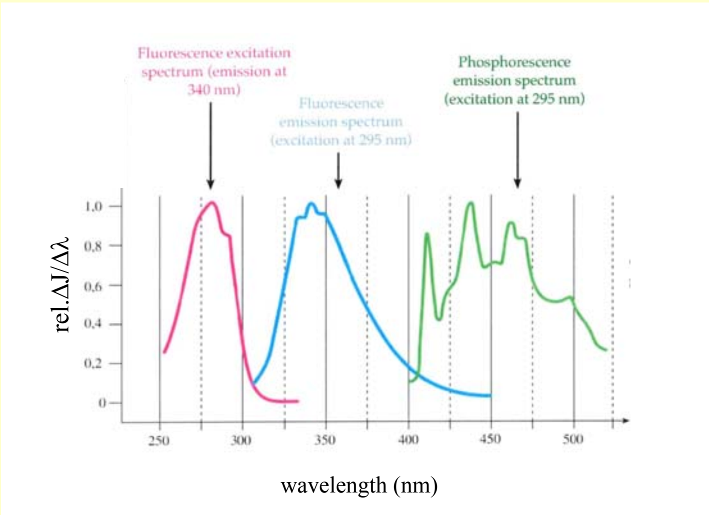
molecules: band spectrum



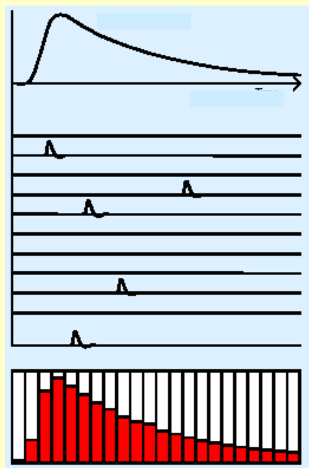
$\lambda_{excitation} \leq \lambda_{fluorescence} < \lambda_{phosphorescence}$

Stokes shift

E.g.: Corresponding spectra of triptophane



Excited-state lifetime

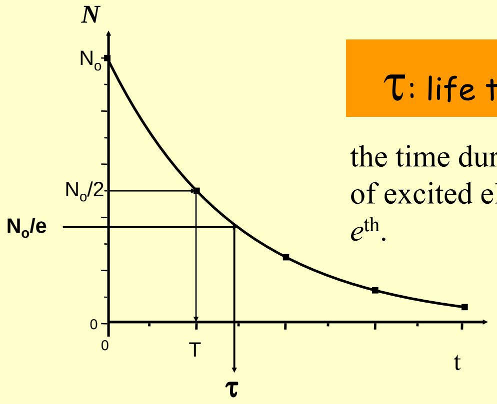


Single photon counting

Measuring of time gap between excitation and photon emission. Statistical analysis of large number of measurements.

Number of excited electrons → $N = N_0 e^{-\frac{t}{\tau}}$ ← time after excitation

- Exponential function

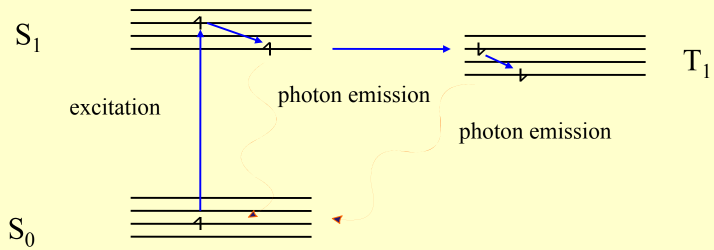


τ: life time

the time during which the number of excited electrons decreases to its e^{th}.

Typical excited-state lifetimes

Lifetime the time during which the number of excited electrons decreases to its e^{th}.



Fluorescence	Phosphorescence
short	long
$10^{-9} - 10^{-7}$ s	$10^{-3} - 10^{2}$ s

Is excitation always followed by photon emission?

Excited state decay can be caused by mechanisms other than photon emission and are therefore often called "non-radiative rates„.

These can include: chemical reaction, dynamic collisional quenching, near-field dipole-dipole interaction, internal conversion and intersystem crossing.

Is excitation always followed by photon emission?

Quantum yield $A \xrightarrow{h\nu} B$

Reciprocal of the number of absorbed photons for one photon emission

Fluorescence quantum yield (Q_F)

$$Q_F = \frac{\text{number.of.photons.emitted}}{\text{numbe.of.photons.absorbed}}$$

$$Q_F \leq 1$$

Types of luminescence

fluorescence

phosphorescence

They can be characterized by

- emission spectrum
  - types
  - position of peaks
  - amplitude
- lifetime
- quantum yield

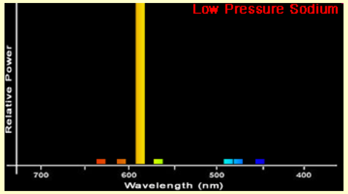
Application fields of luminescence

Light sources (lightning, fertilization, sunbeds, photomedicine...)

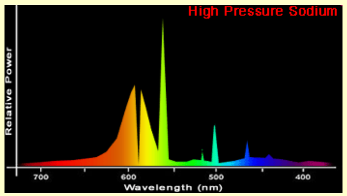
- concentration determination (flame photometer)
- luminescence spectroscopy
- luminescence microscopy
- dosimetry (see later)
- archeology
- architecture
- safety controll … many more

Luminescent light sources

Metal vapor lamps

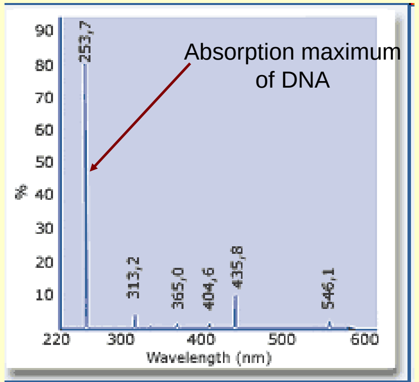


Low-pressure Na-vapor lamp



High-pressure Na-vapor lamp

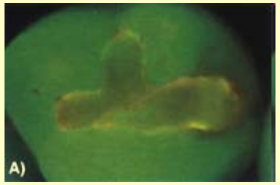
Low-pressure Hg-vapor lamp



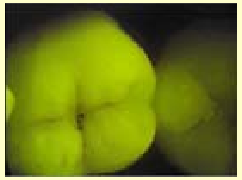
Emission spectrum

„germicid lamp"

Application in dental medicine

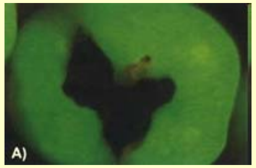


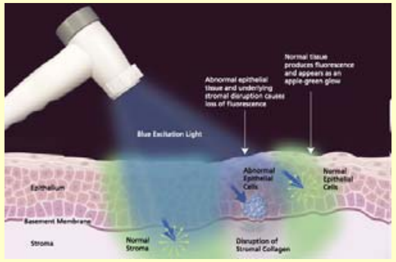
Red fluorescence indicates the activity of identifies cariogenic bacteria



Auto-fluorescence of teeth. When teeth are illuminated with high intensity blue light they will start to emit light in the green part of the spectrum.

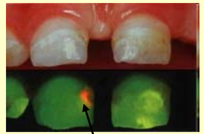
amalgam restoration





Healthy and malignant tissues different fluorescent properties

Teeth native and fluroescent images

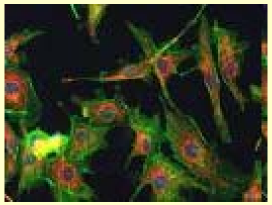


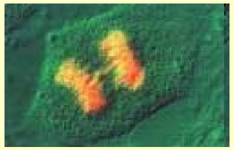
Active caries

caries

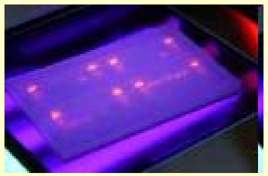
Tooth native and fluroescent image

Luminescent microscopy





Laboratory application in many ways



And more...

Damjanovich, Fidy, Szöllősi: Medical Biophysics

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- 2.2.4
- 2.2.6

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- 3.3.2 pp. 411-413
- 3.3.3