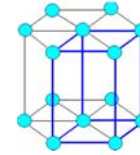


Physical bases of dental material science

Irén Bárdos-Nagy

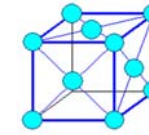
Submicroscopic structure of metals



hexagonal

Ti, Cd, Co, Zn

Space filling factor 74 %



face-centered cubic

Ag, Au, Pd, Pt, Al, Cu, Ni

Space filling factor 74 %

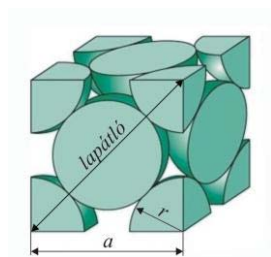


body-centered cubic

Fe, Cr

Space filling factor 68 %

The Ni has face centered cubic crystal structure. Calculate the space filling factor for this metal. How large is the density of the Ni, if the atomic mass is 58.7 and the atomic radius is 0.125nm?



In one unit cell there are 6 * 1/2 atoms and 8 * 1/8 atoms

↓
4 atoms

$$V_a = \frac{4}{3} r^3 \cdot \pi$$

$$a \cdot \sqrt{2} = 4r \rightarrow a = 2 \cdot r \cdot \sqrt{2}$$

$$V_{cell} = 8 \cdot \sqrt{8} \cdot r^3$$

$$\frac{4 \cdot V_a}{V_{cell}} = \frac{4 \cdot \frac{4}{3} \cdot r^3 \cdot \pi}{8 \cdot \sqrt{8} \cdot r^3} = 0,74$$

$$\rho = \frac{m}{V} = \frac{4 \cdot \frac{M}{N_A}}{V_{cell}} = \frac{4 \cdot \frac{58,7}{6 \cdot 10^{23}}}{8 \cdot \sqrt{8} \cdot (0,125 \cdot 10^{-7})^3} = 8,85 \left(\frac{g}{cm^3} \right)$$

Alloys

partial or complete solid solutions of one or more elements in a metallic matrix

metal + metal (Fe+Cr+Ni, Au+Cu)

metal + non metal (Fe+C)



The aim: to modify (to improve) the properties

- hardness and rigidity (Au + Cu)
- tensile strength
- shear strength
- to avoid or reduce the corrosion (Fe, Co, Ni, + Cr)
- to increase the adhesion on metal-ceramic surfaces (precious metal+Fe, Sn, In)



Determination of composition

$$\text{weight \% : } c_{m1} = \frac{m_1}{m_1 + m_2} \cdot 100(\%)$$

properties!!

$$\text{molar \% : } c_{v1} = \frac{v_1}{v_1 + v_2} \cdot 100(\%)$$

$$c_{v1} = \frac{\frac{m_1}{M_1}}{\frac{m_1}{M_1} + \frac{m_2}{M_2}} \cdot 100 = \frac{v_1}{v_1 + v_2} \cdot 100 = \frac{m_1 \cdot M_2}{m_1 \cdot M_2 + m_2 \cdot M_1} \cdot 100(\%)$$

$$c_{m1} = \frac{m_1}{m_1 + m_2} \cdot 100 = \frac{v_1 \cdot M_1}{v_1 \cdot M_1 + v_2 \cdot M_2} \cdot 100 = \frac{c_{v1} \cdot M_1}{c_{v1} \cdot M_1 + c_{v2} \cdot M_2} \cdot 100(\%)$$

The molar ratio in a gold – silver alloy is 2. How large is the molar percent?
How large is the mass percent of the two components in this metal?

$$M_{Au} = 197g \quad M_{Ag} = 108g$$

$$c_{vAg} = \frac{1}{1+2} \cdot 100(\%) = 33,3\%$$

$$c_{vAu} = \frac{2}{1+2} \cdot 100(\%) = 66,7\%$$

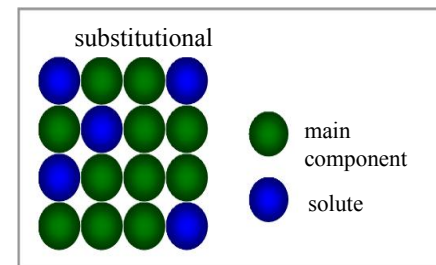
$$c_{mAg} = \frac{m_{Ag}}{m_{Ag} + m_{Au}} \cdot 100(\%) = \frac{0,333 \cdot 108}{0,333 \cdot 108 + 0,667 \cdot 197} \cdot 100 = 21,5\%$$

$$c_{mAu} = \frac{m_{Au}}{m_{Ag} + m_{Au}} \cdot 100(\%) = \frac{0,667 \cdot 197}{0,333 \cdot 108 + 0,667 \cdot 197} \cdot 100 = 78,5\%$$

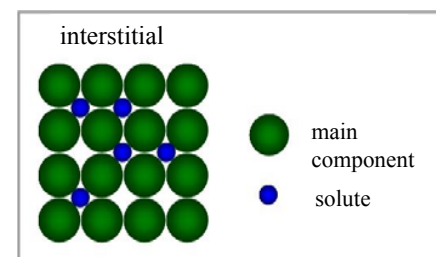
Classification:

- according the application (inlay, corona of teeth)
- on the base of the main component (Au, Pd, Pt, Fe)
- on the base of the number of components (biner, terner, quaterner)
- on the base of the main 3 components (Au-Pd-Ag, Ni-Cr-Be)
- on the base of the phase diagram (solid solution, eutectic alloy, peritectic alloy, metal compound)

Solid solutions



examples: Cu-Ni, Pd-Ag, Au-Cu, ...



examples: Fe-C, CP Ti (O, C, N, H), ...

homogeneous structure

criteria of formation

similar atomic radii (less then 15% diff.)
same crystal structure
similar electronegativities
similar valency

the atomic radius of the solute is smaller
the amount of the solute is less than 10%

properties of solid solutions

flexibility changes

strength increases

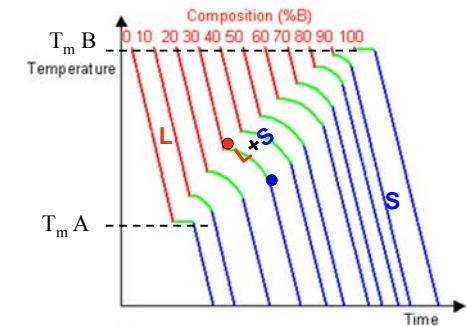
hardness increases

ductility changes

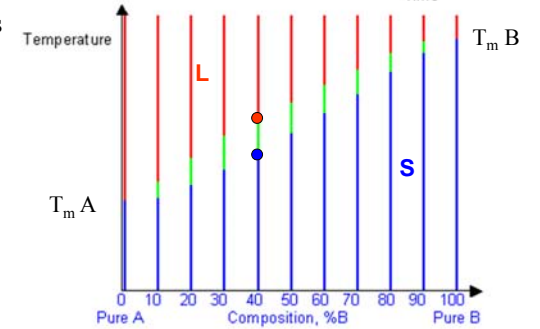
plasticity decreases

metal	atomic radius (nm)	lattice	electro-negativity
Au	0,2882	fcc	2,4
Pt	0,2775	fcc	2,2
Pd	0,2750	fcc	2,2
Ag	0,2888	fcc	1,9
Cu	0,2556	fcc	1,9
Ni	0,25	fcc	1,8
Sn	0,3016	tetragonal	1,8

cooling curve of solid solutions

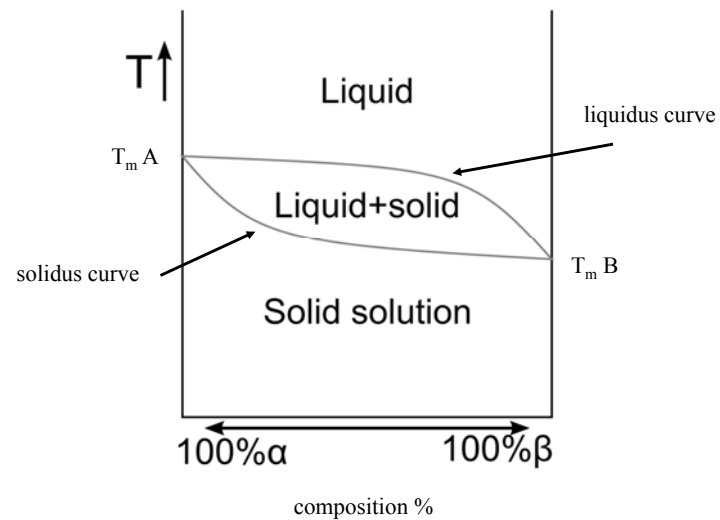


phase diagram of solid solutions

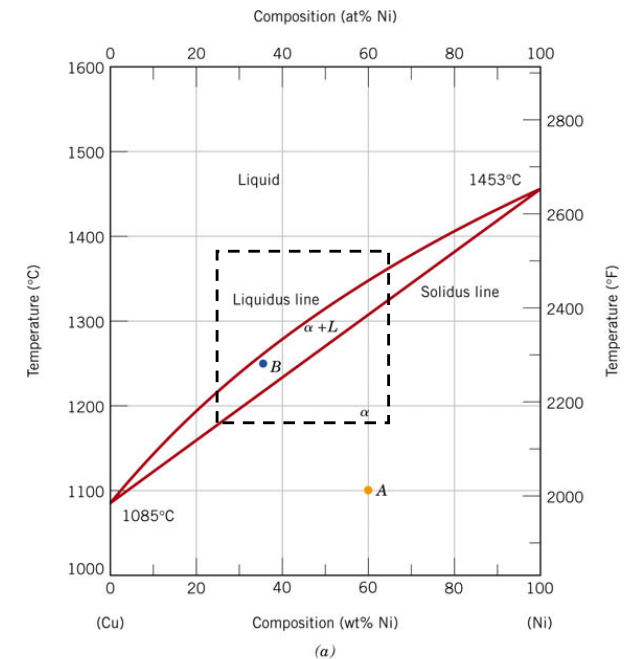


equilibrium !!

equilibrium

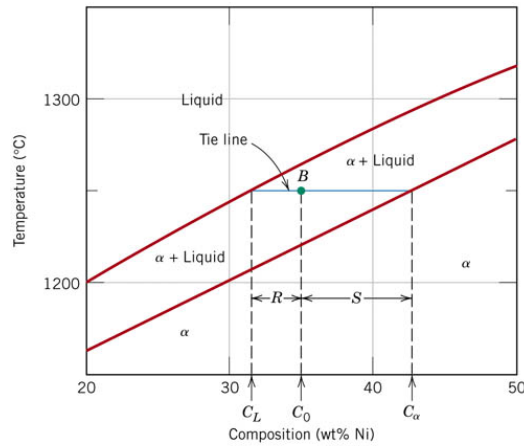


example:
Cu + Ni



Calculation of the composition and the ratio of the different phases

what is the composition at the B point



Liquid phase composition:
31,5 wt % Ni + 68,5 wt % Cu

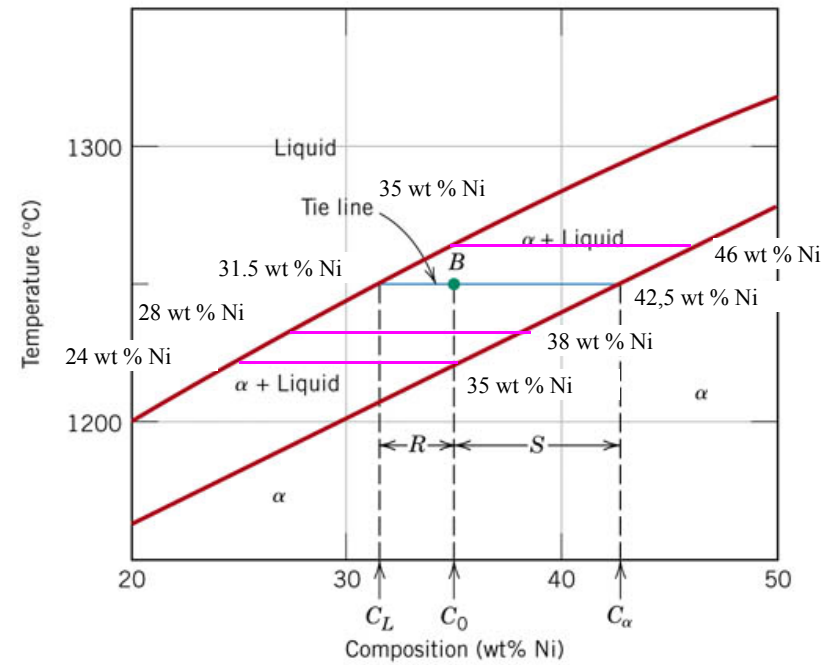
Solid phase composition:
42,5 wt % Ni + 57,5 wt % Cu

Liquid phase ratio:

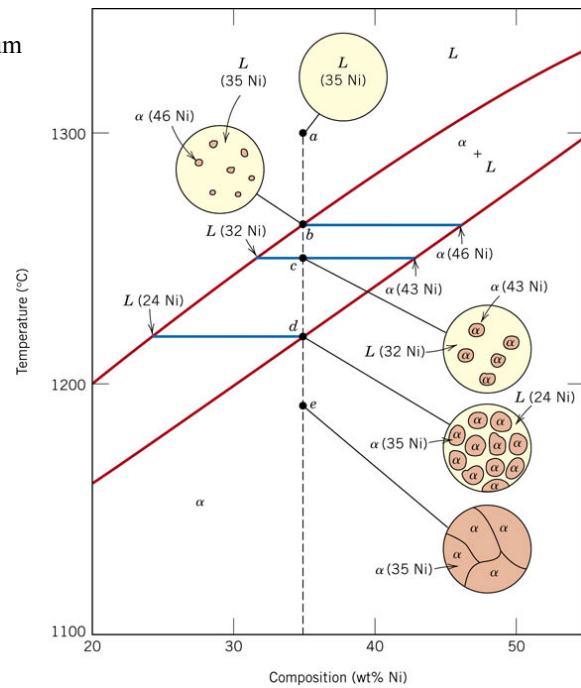
$$\frac{S}{R + S} = 68\%$$

Solid phase ratio:

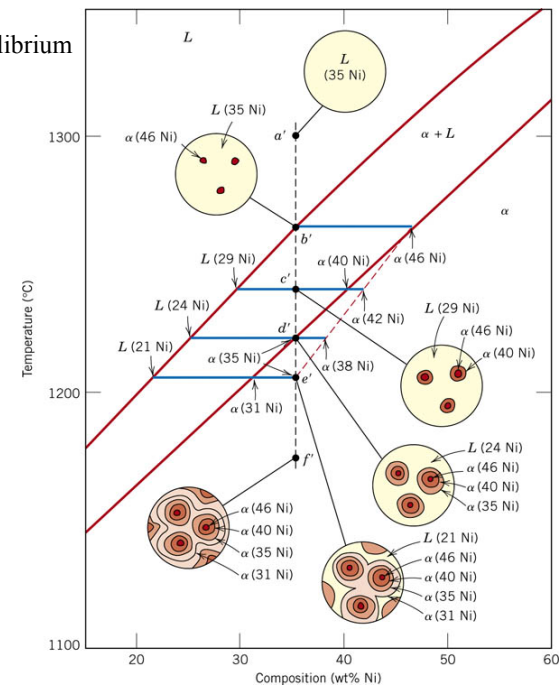
$$\frac{R}{R + S} = 32\%$$



equilibrium

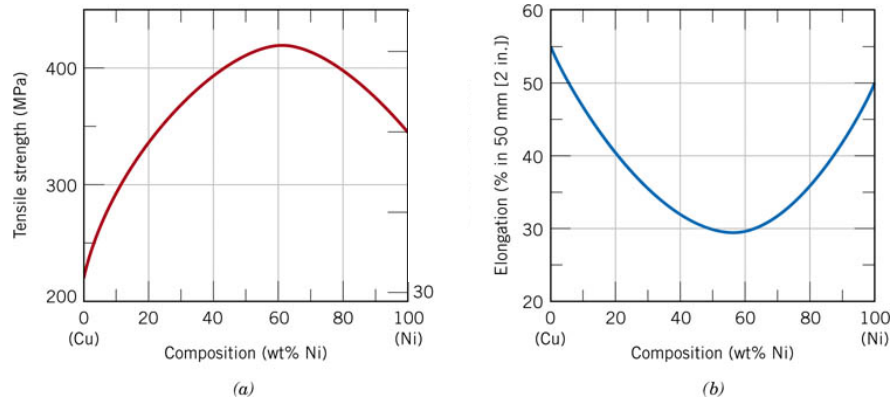


non equilibrium cooling



segregation
seedy structure
heterogeneous
↓
homogenization

Influence of the solute material on different physical properties of alloys



Metal compounds

Definite stoichiometry

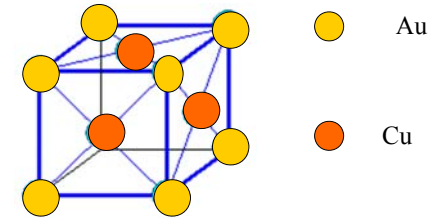
example: Au-Cu

50 wt %Au-50 wt %Cu

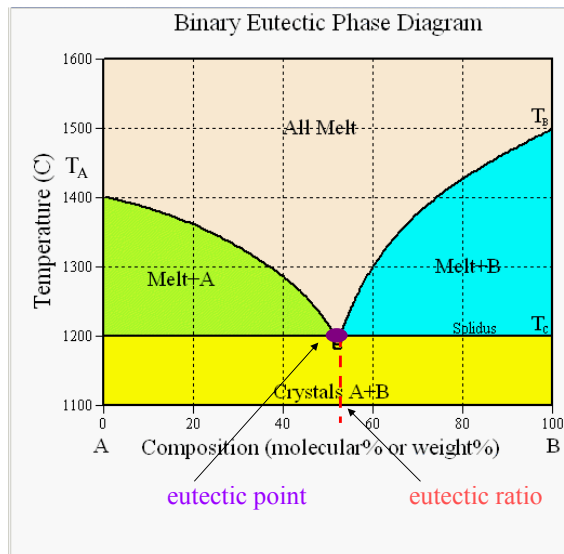
< 400 C



in the amalgam: Ag_3Sn
 Cu_6Sn_5



Eutectic alloys



insolubility in the solid phase

pure metal crystallites

heterogenous structure

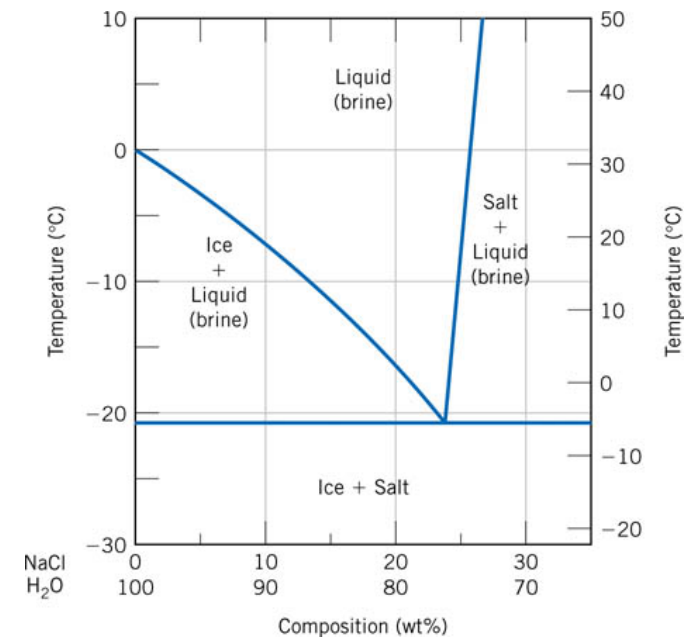
examples

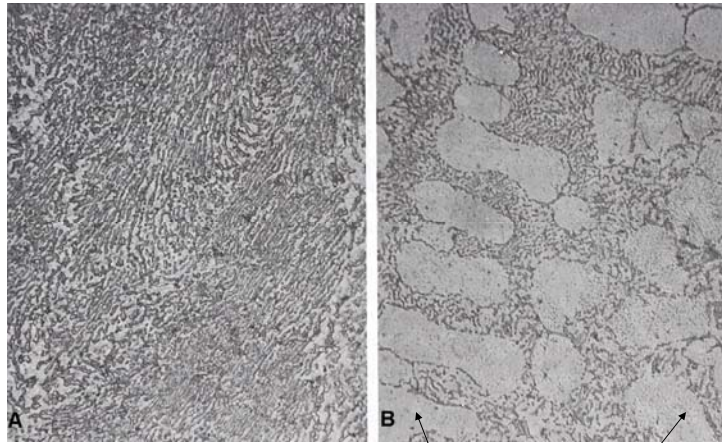
77% H_2O + 23 % NaCl

$T_E = -21^\circ\text{C}$

Wood-metal (Bi-Pb-Cd-Sn)

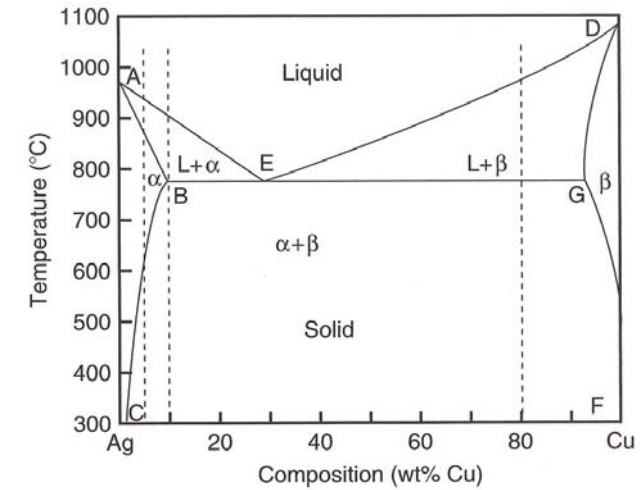
$T_E = 68^\circ\text{C}$





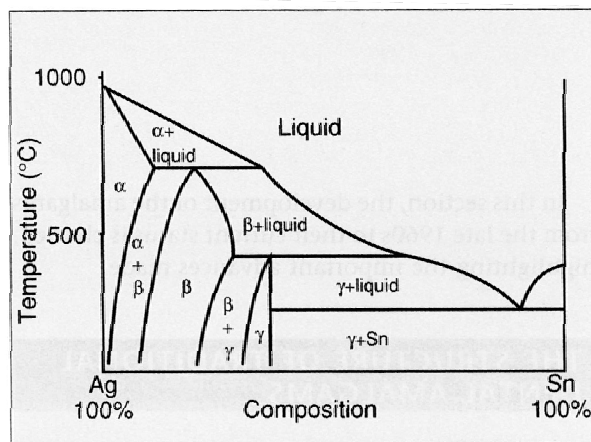
62%Sn-38%Pb
eutectic alloy

75%Sn-25%Pb
(Sn rich islands)



dental amalgam

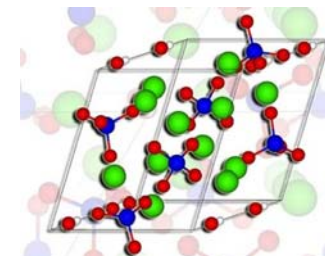
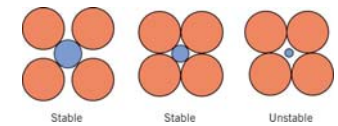
composition	
metal	% (wt/wt)
Hg	50
Ag	34
Sn	13
Cu	2
Zn	1



γ phase: Ag₃Sn

Ceramics: compounds of metal and non-metal elements

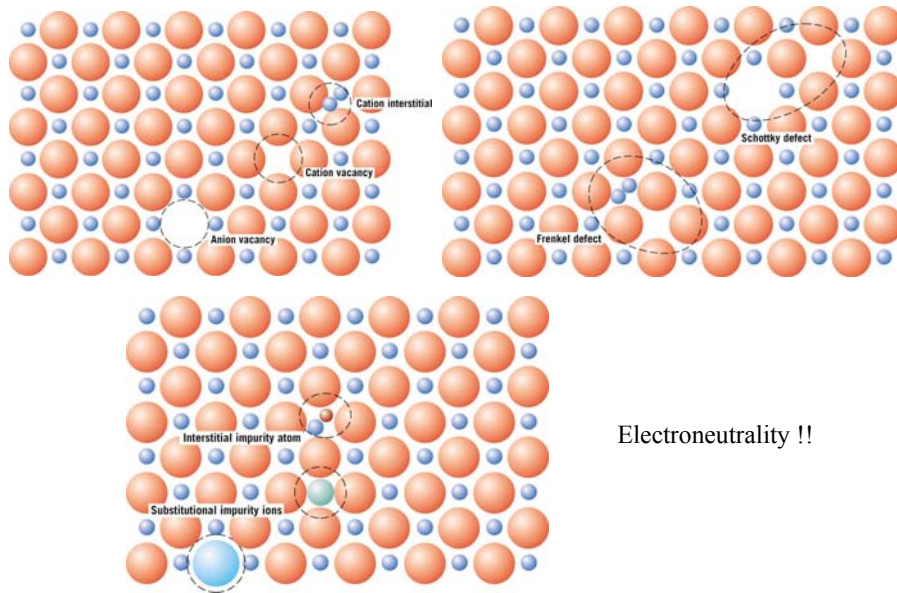
- mainly ionic bond between the components
- the + ions are smaller
- crystalline or amorphous structure
- phase diagram like in the case of metals



apatite

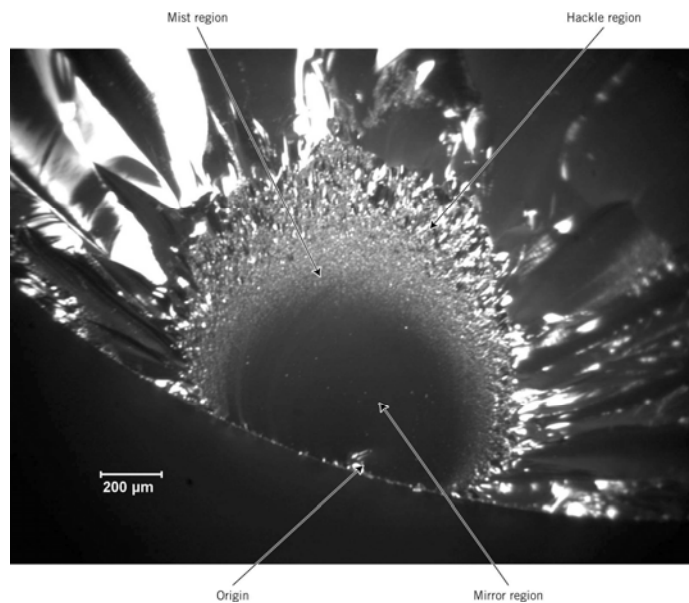


Defects of ceramics

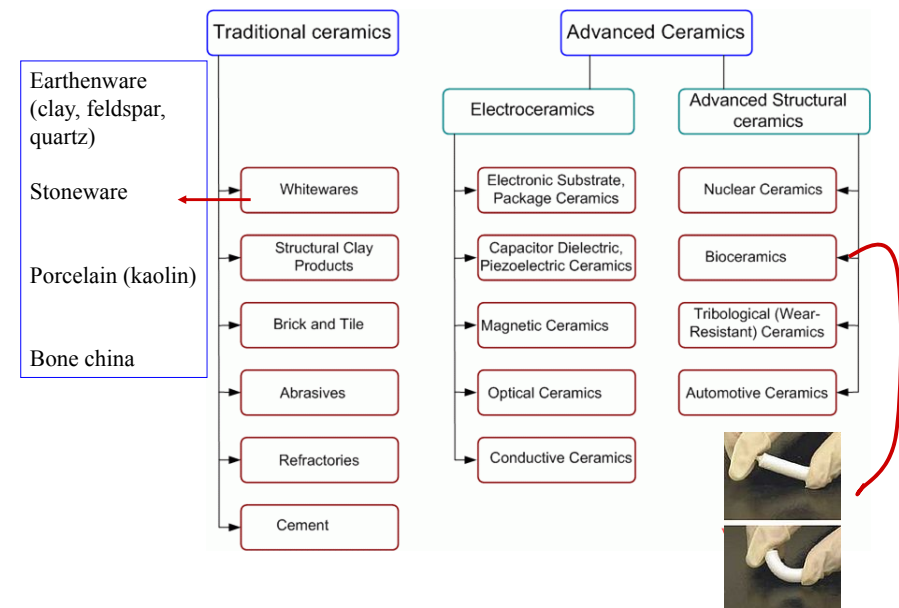


General properties of ceramics

- solid at room temperature
- fragility
- large hardness and rigidity
- strong in compression
- weak in tension and shearing
- poor heat-shock tolerance
- proof against corrosion and heat
- not conduct electric current (insulator)
- biocompatibility



Application base classification of ceramics





← Earthenware pottery



Porcelain →

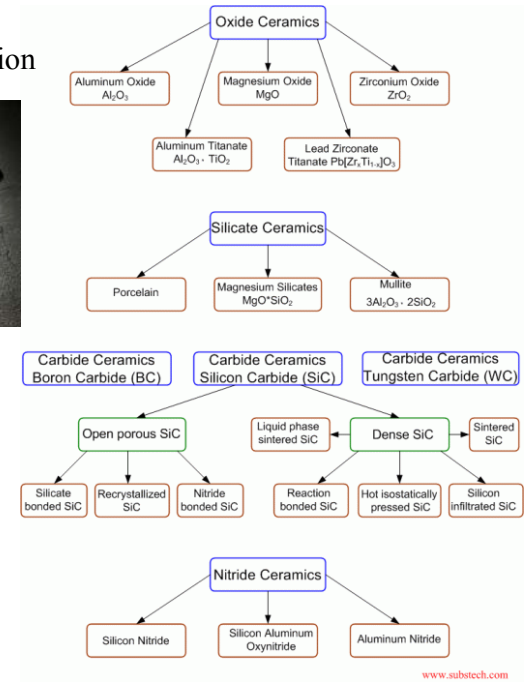
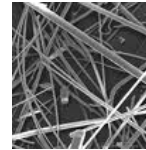


← Stoneware vase

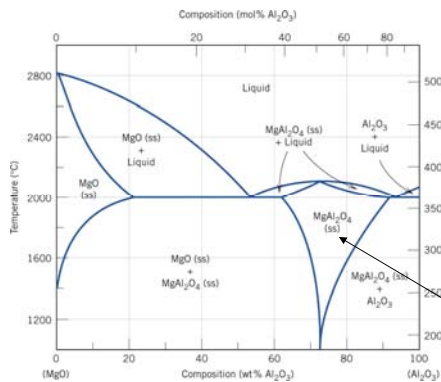
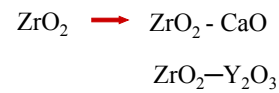


Bone china →

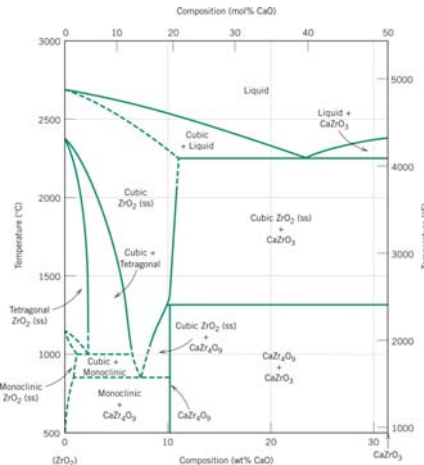
composition base classification of ceramics



Oxide ceramics

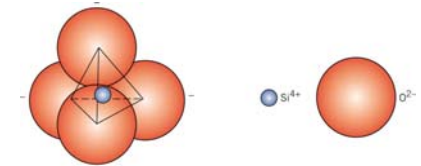
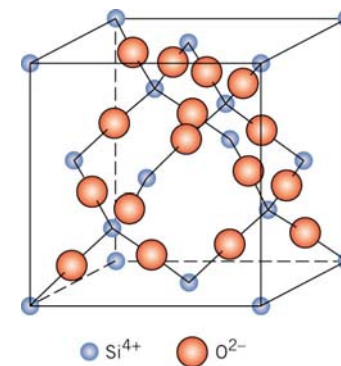


MgO - Al₂O₃
(MgAl₂O₄ spinel)



Silicate ceramics

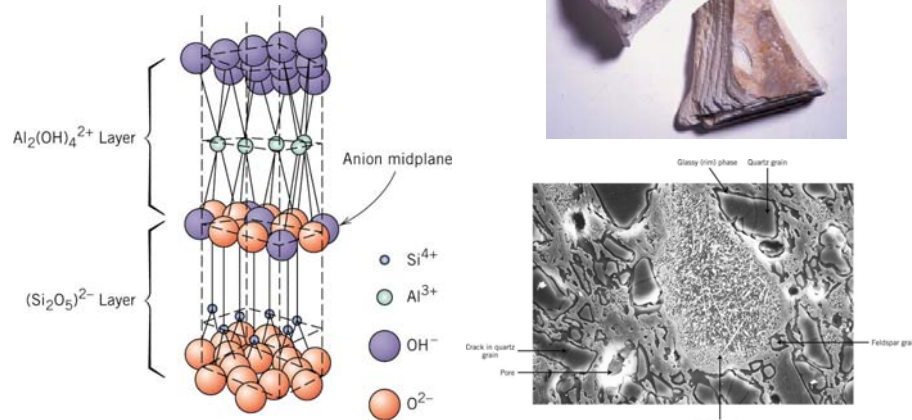
crystal structure of silica



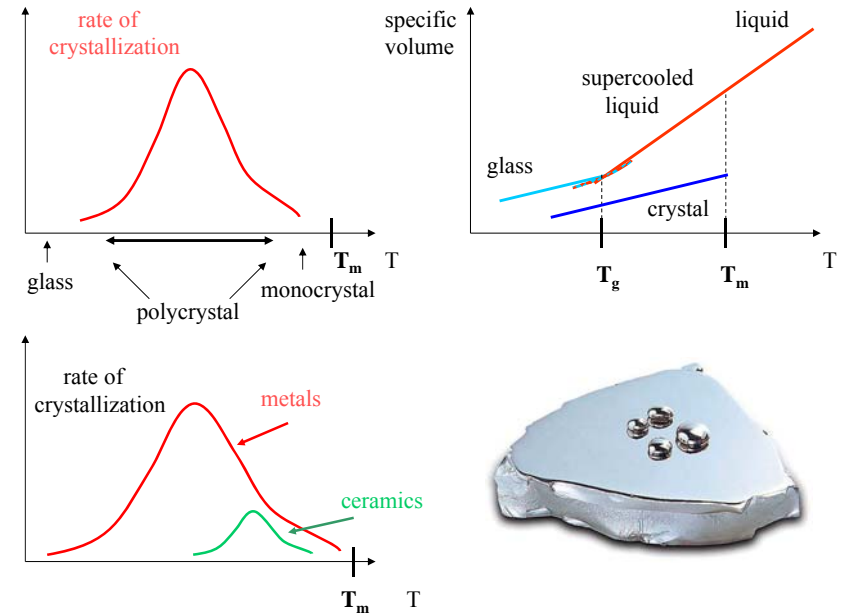
amorphous sodium silicate

Porcelain

crystal structure of kaolin



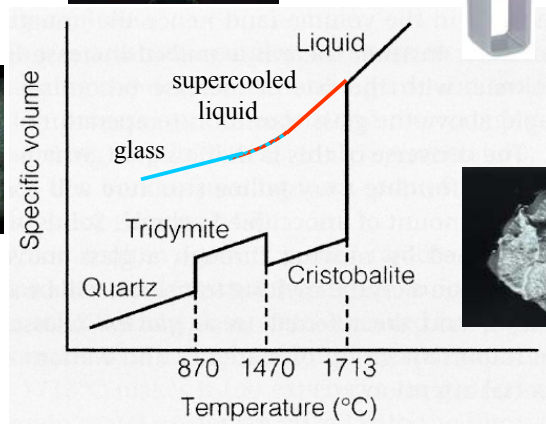
Crystallization ↔ glass formation



SiO₂



silica glass



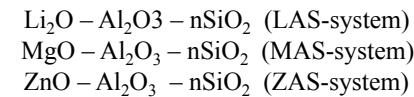
Glass ceramic

share many properties with both glasses and ceramics

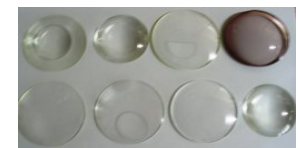
glass ceramics structure:

they have amorphous phase
+ crystalline one (controlled crystallization)

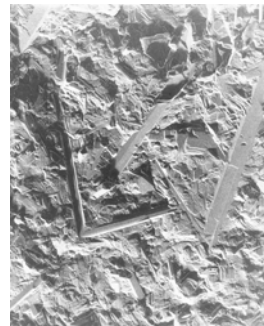
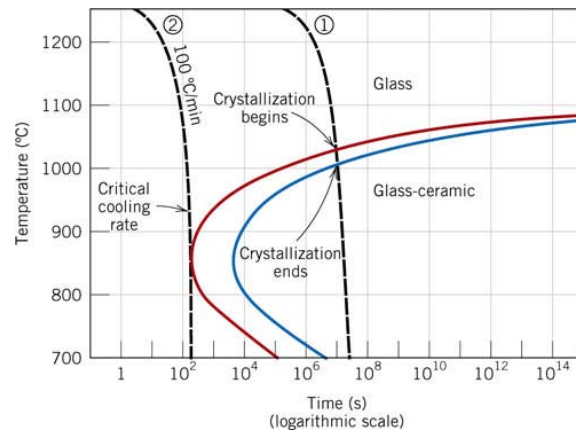
main components:



- properties:
- mechanically strong
 - no pores inside the structure
 - low heat conduction coefficient
 - quick temperature changes (up to 800 -1000 C°)
 - coefficient of thermal expansion can vary with phase ratios (it can be negative or even zero)



Formation of glass ceramics



Dental application of ceramics

- crowns, veneers
- bridges
- fillings
- dental implants
- orthodontic brackets
- cements
- polishing materials

