

Medical Biophysics I.

Interactions of light with matter.

Light scattering, light absorption. Basic principles of absorption spectrometry.

(Help to understand and learn exam question 12. and to measurement „Light absorption“.)

8th Lecture, 25th October 2016

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12th exam question

How the light can interact with atoms and molecules?

Basic concepts: Light scattering, Rayleigh-scattering, Mie-scattering, light absorption.

Phenomena: Blue sky, white and grey clouds, transparency.

Physical quantities: Light intensity (J), electric dipole momentum (p), absorbance (A).

Laws, relations and the way leading to them: Attenuation of light intensity passing through medium. Dependence of scattered light intensity on the wavelength. Lambert–Beer-law.

Applications: Measurement of static light scattering and absorption, absorption spectrometry, determination of concentration. Dark field microscope.

Repetition I. - Light

What is light? – visible electromagnetic wave

Wave: quantity(ies) changing periodically in SPACE and TIME

Electromagnetic wave:

electric field (E) and magnetic flux density („field” (B)) -||-

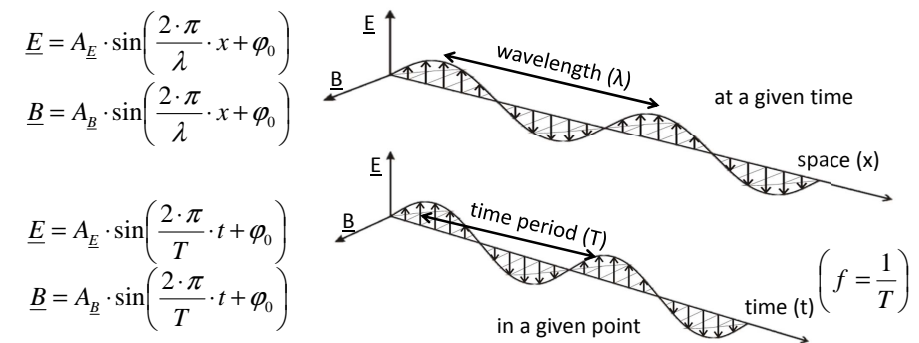
$$\vec{F} = q \cdot \vec{E}$$

$$\vec{F}_L = q \cdot \vec{v} \times \vec{B}$$

Electric field is defined as the electric force per unit charge

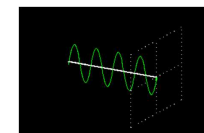
Magnetic fields are produced by electric currents (moving charges), which can be macroscopic currents in wires, or microscopic currents associated with electrons in atomic orbits

Repetition I. - Light

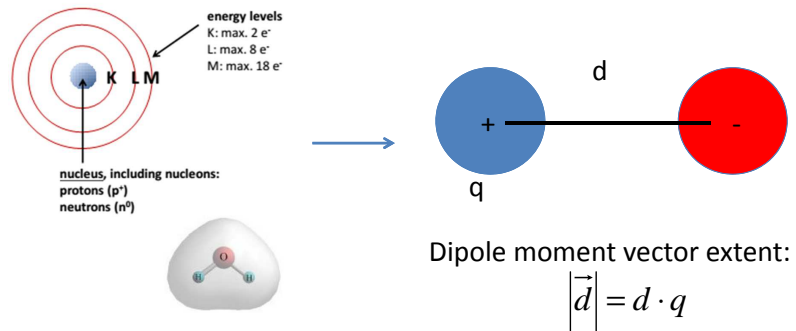


TIME and SPACE: $c = \lambda \cdot f$

$$\varepsilon = h \cdot f \quad E = \sum N \cdot \varepsilon_{\lambda} \quad J = \frac{E}{A \cdot t}$$



Repetition II. Matter: Atoms and Molecules

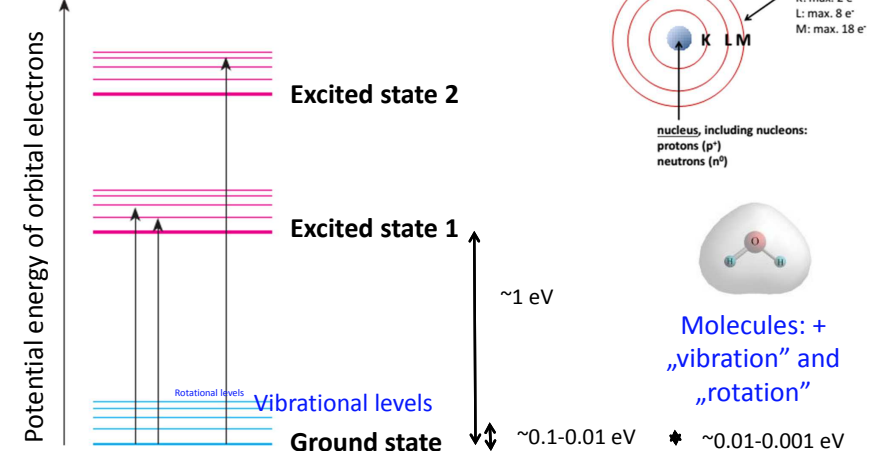


Polarizability (α): extent of a dipole that is generated by a unit of electric force-field

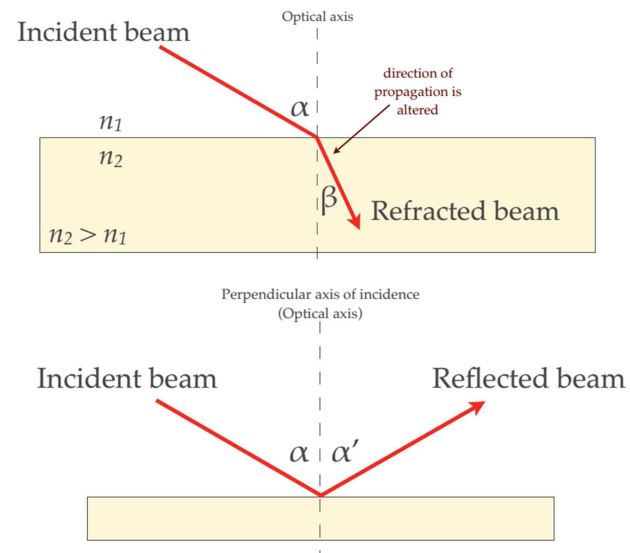
$$\vec{d} = \alpha \cdot \vec{E}$$

Repetition II. Atoms and Molecules

Electronic and vibrational energy
levels of atoms and molecules



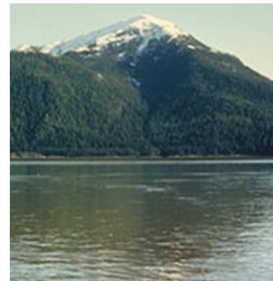
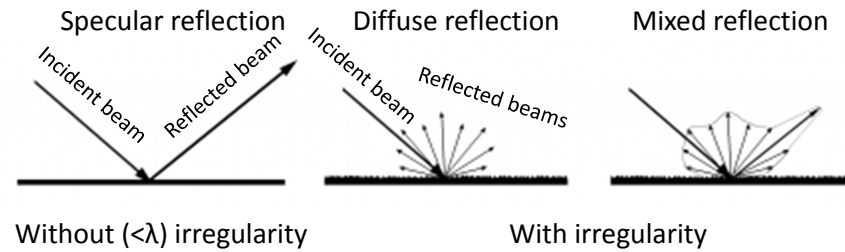
InteractionS – repetition



InteractionS – repetition

Experiment 1.

InteractionS – Reflection



InteractionS – Reflection

Coefficient of reflection

$$\rho_{\lambda} = \frac{J_{\text{reflected}}(\lambda)}{J_{\text{incident}}(\lambda)}$$

depends on:
wavelength
incident angle
material (refractive indices ...)

$$\rho_{\lambda} = \frac{J_{\text{reflected}}(\lambda)}{J_{\text{incident}}(\lambda)} = \left(\frac{n_2 - n_1}{n_2 + n_1} \right)^2$$

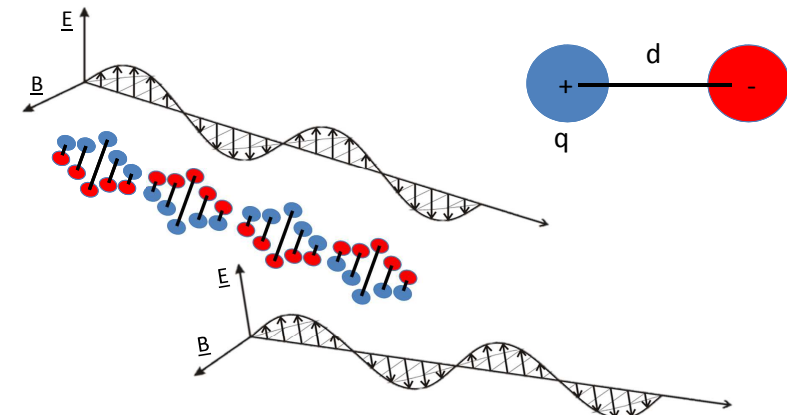
equation is true **IF** incident angle is 0° ,
material is transparent (see later)

InteractionS

Experiment 2.

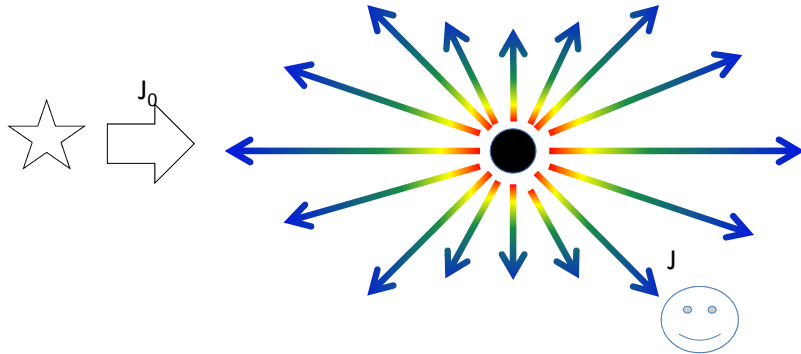
InteractionS – Scattering

1. Elastic scattering (f does NOT change)



InteractionS – Rayleigh Scattering

Small, individual particles – individual oscillations without interference

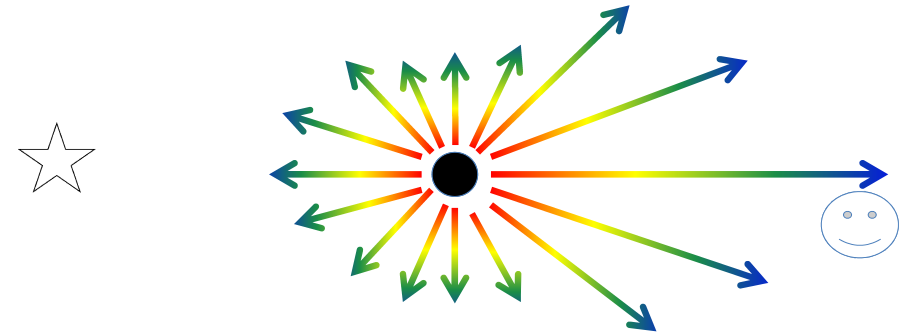


$$J_{\lambda} = J_0 \cdot \frac{8 \cdot \pi^4 \cdot N \cdot \alpha^2}{\lambda^4 R^2} \cdot (1 + \cos^2(\beta))$$

N: number of scatterers
 α : polarizability
 λ : wavelength
 R: distance from scatterer
 β : observing angle

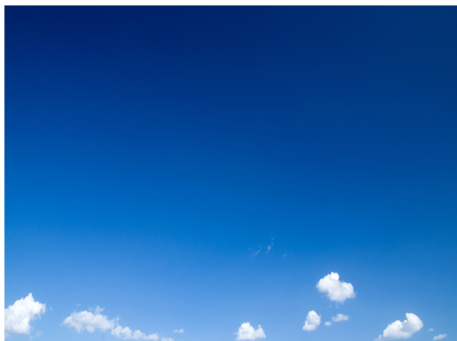
InteractionS – Mie Scattering

Large particles, group of atoms, molecules (size $> \lambda/10$)
 – constructive interference



wavelength independent
 large forward scattering
 Size dependent:
 Size increasing: intensity increasing
 Size increasing further ($> \lambda \cdot 10$): intensity decreasing

InteractionS – Rayleigh Scattering



Blue sky
 Red sunset



InteractionS - Mie scattering



Glare, white cloud



Gray cloud

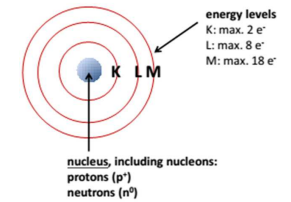
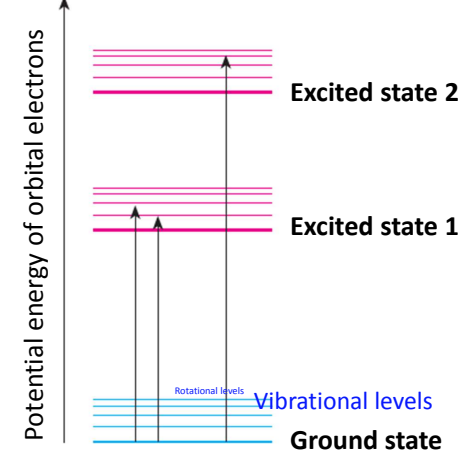
InteractionS

Experiment 3.

InteractionS - Absorption

Resonance: induced oscillation – energy levels

Electronic and vibrational energy levels of atoms and molecules



Molecules: + „vibration” and „rotation”

InteractionS - Scattering

2. Inelastic scattering (f changes)

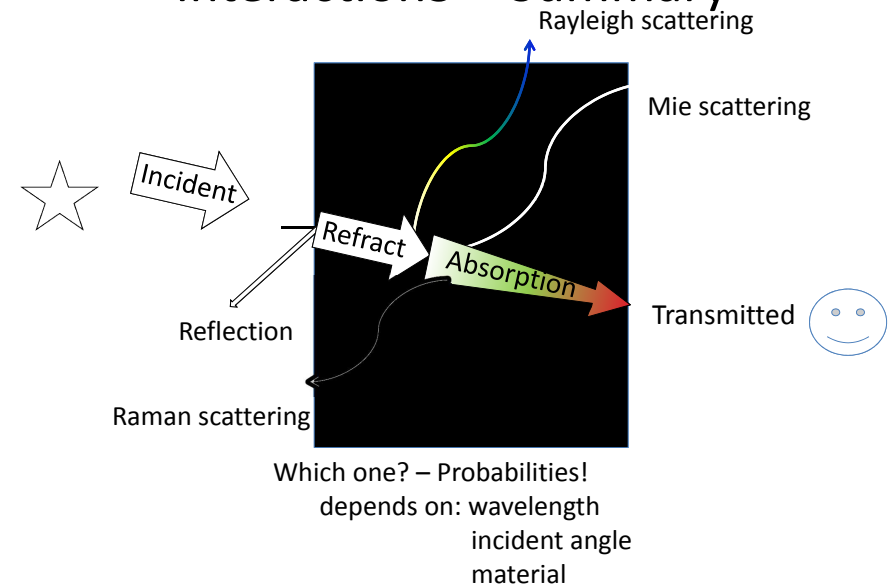
- energy absorption between vibrational levels + re-emission
- short life time in the excited state (\sim fs)

Raman scattering

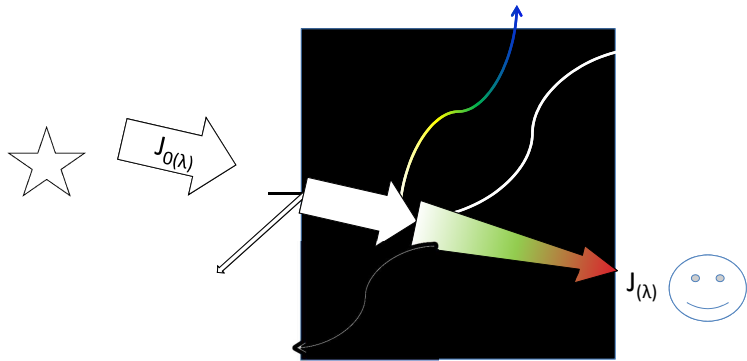
Stokes: scattered ϵ reduced

Anti-Stokes: scattered ϵ increased

InteractionS - Summary



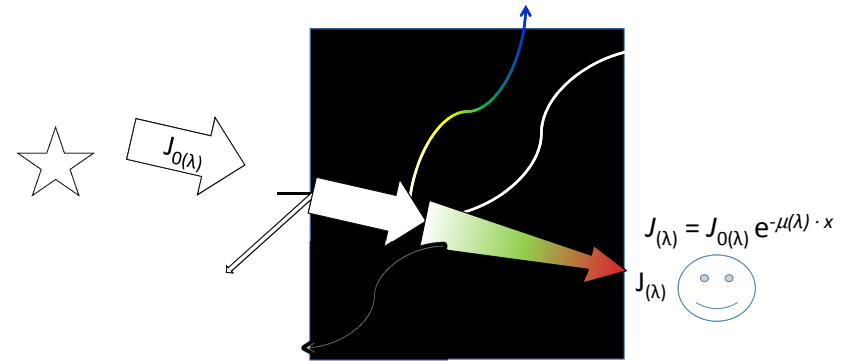
Intensity attenuation - Decay law



$$J_{(\lambda)} = J_{0(\lambda)} e^{-\mu(\lambda) \cdot x}$$

J: intensity
 μ : attenuation coefficient
 x: layer thickness

Transmittance, Optical density



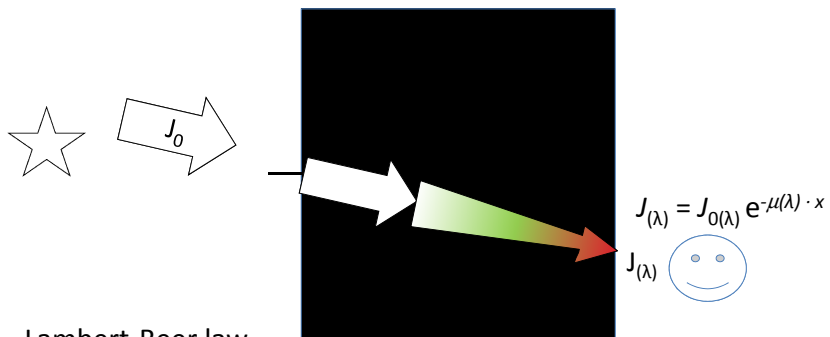
$$J_{(\lambda)} = J_{0(\lambda)} e^{-\mu(\lambda) \cdot x}$$

$$T = \frac{J_{(\lambda)}}{J_{0(\lambda)}} \quad (\%)$$

T: transmittance
 OD: optical density

$$OD = \log \frac{J_{0(\lambda)}}{J_{(\lambda)}} = \log e \cdot \mu_{(\lambda)} \cdot x$$

Absorption



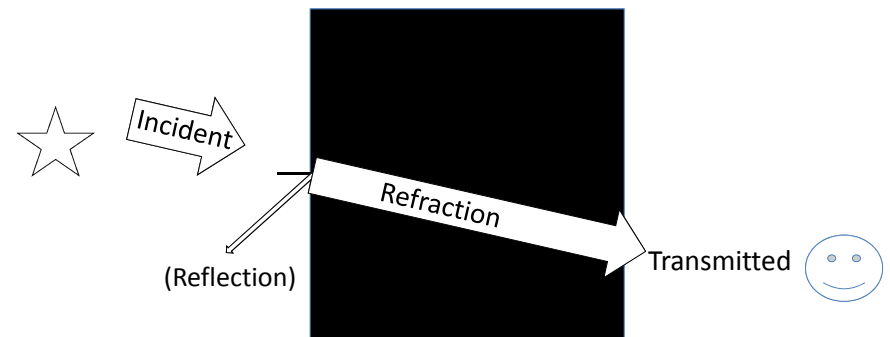
Lambert-Beer law

$$A = \log \frac{J_{0(\lambda)}}{J_{(\lambda)}} = \log e \cdot \mu_{(\lambda)} \cdot x = \epsilon_{(\lambda)} \cdot c \cdot x$$

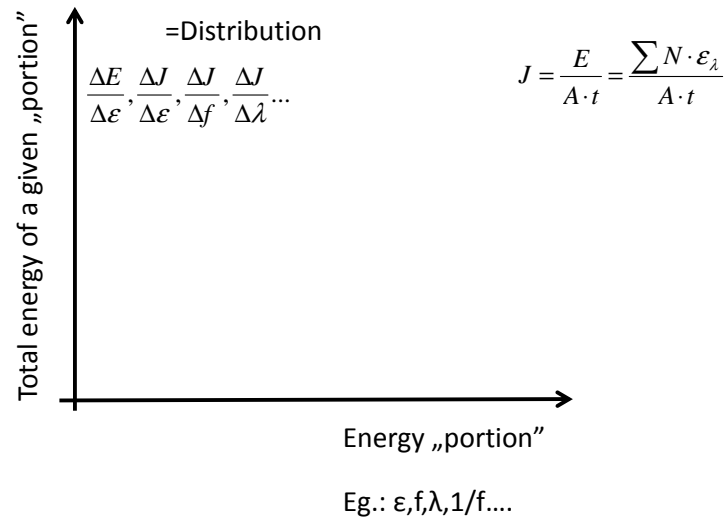
Conditions!!!

μ : absorption coefficient
 ϵ : (decadic) molar extinction coefficient
 c: concentration

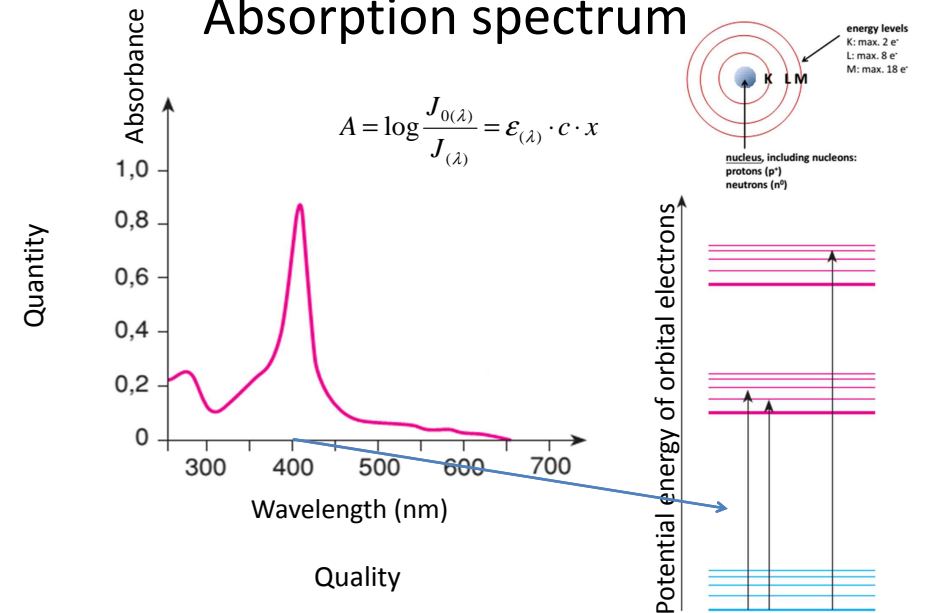
Transparency



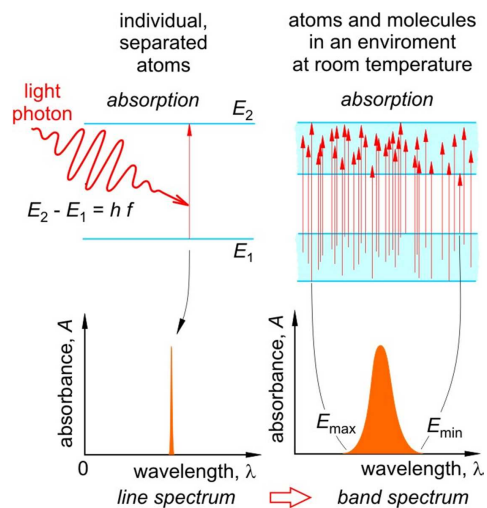
Spectrum



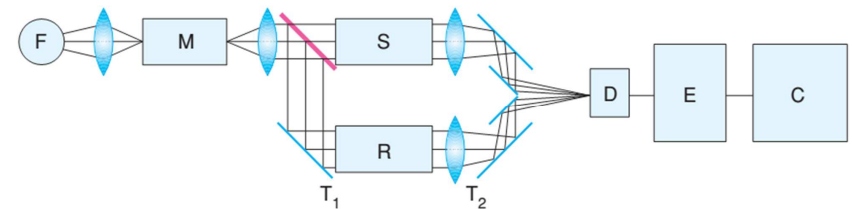
Absorption spectrum



Spectrum broadening



Absorption spectrum measurement



Monocromators:

Prism: refraction – advantage: total intensity

Optical grating: diffraction (interference image) – advantage: linear („ghosts”: harmonics)

Scattering measurements

Concentration measurements:

Turbidimetry: (high intensity of scattering)
measuring OD (condition: no absorption)

Nephelometry: (low intensity of scattering)
collecting scattered light in a given angle

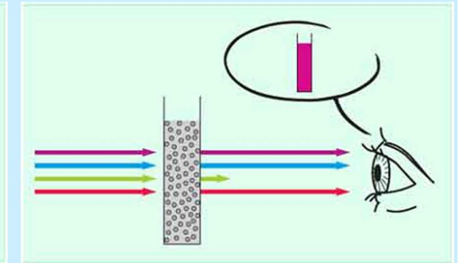
Particle size and geometry measurement:

Dynamic light scattering: (size)
collecting scattered light in a given angle
scattered intensity changes in time

Static light scattering: (size, geometry)
collecting scattered light in a given angle range (multiple angle)

Color - absorption

Complementary colors



The color of matter is related to its **absorptivity**, reflectivity and light scattering. The human eye sees the complementary color to that which is absorbed

Colors



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