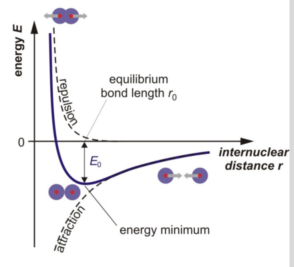
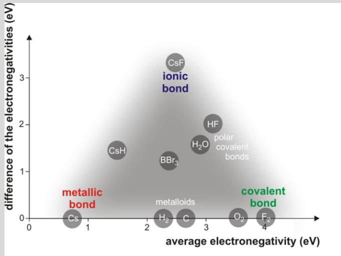
Repetition

Atomic interactions

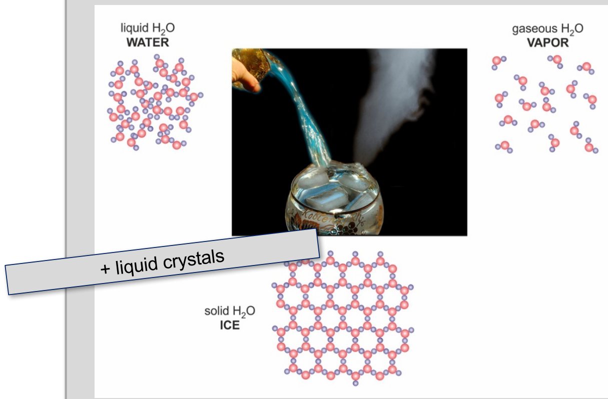


Bond types



2

States of matter

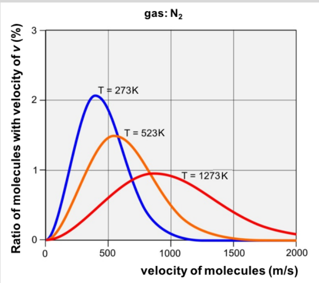


3

Gases

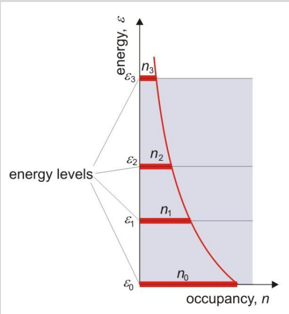
Interpretation of temperature:

$$\frac{1}{2}m\overline{v^2} = \frac{3}{2}kT$$



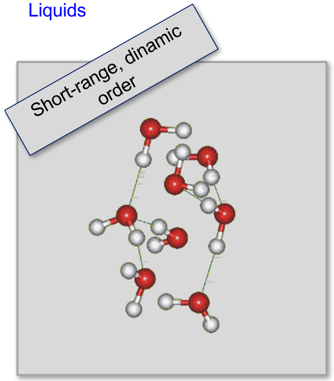
Boltzmann-distribution:

$$n_i = n_0 \cdot e^{-\frac{\varepsilon_i - \varepsilon_0}{kT}}$$



4

Liquids

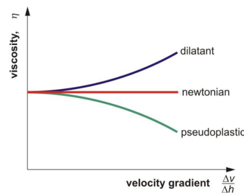


Viscosity:

$$F_s = \eta \cdot A \cdot \frac{\Delta v}{\Delta h}$$

viscosity (coeficient internal friction)

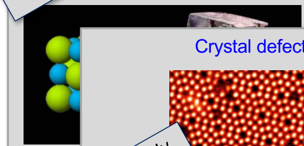
depends on:
- temperature
- velocity gradient
- time of mechanical stress



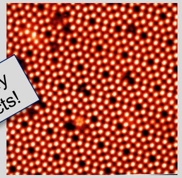
5

Solid - Crystals

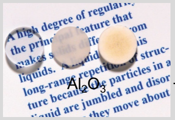
Long-range order



Crystal defects:



Properties strongly depend on defects!



+ Cr^{3+}

Apatite:

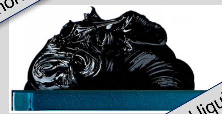
$Ca_{10}(PO_4)_6(X)_2$

entin and bone:
0 nm x 6 nm crystals

namel:
1000 nm x 30 nm crystals

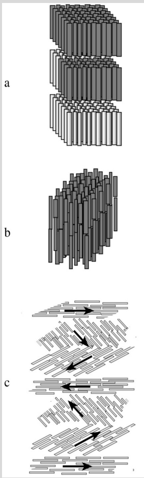
Solid- amorphous

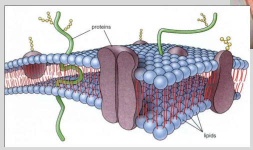
short-range order

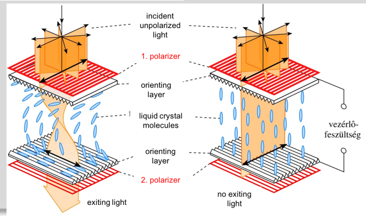


overcooled liquid, glass

Liquid crystals







Physical bases of dental material science 3.

Structure of matter

Interfacial phenomena

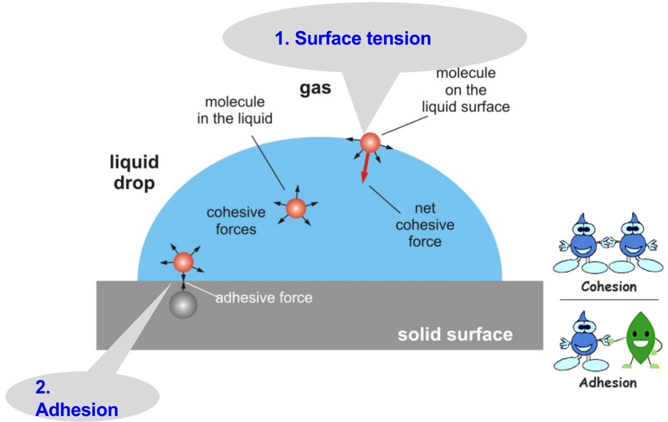
Phase diagram, phase transitions

Highlights:
- ❖ surface tension
- ❖ adhesion – wetting
- ❖ phase – phase diagram
- ❖ phase transition

E-book chapters: 6, 7

Problems: Chapter1: 24, 25, 27, 28, 31

Interfacial phenomena



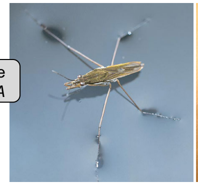
1. Surface tension

Surface tension or specific surface energy (σ):

change in energy with the increase of surface by ΔA

$$\sigma = \frac{\Delta E}{\Delta A} \quad \left(\frac{J}{m^2} = \frac{N}{m} \right)$$

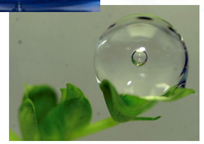
increase in surface



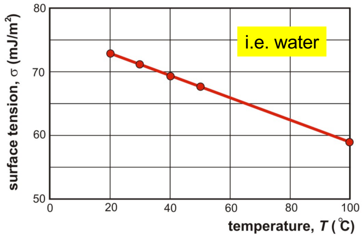

anyag	σ (J/m²)
water	0,073
blood	0,06
saliva	0,05
paraffin	0,025
alcohol	0,023
dentin	0,092
enamel	0,087
mercury	0,484

* in air, 20°C



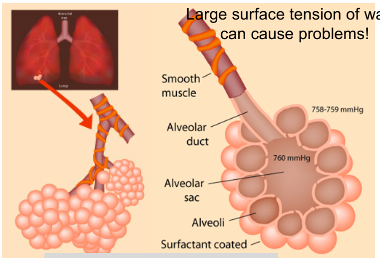



Temperature dependence of surface tension:



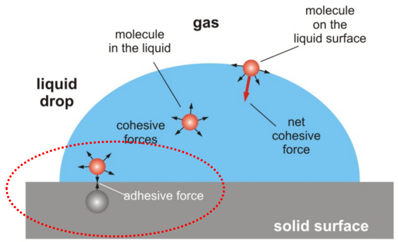
Consequences:

Large surface tension of water can cause problems!



respiratory distress syndrome

2. Adhesion



Types:

- chemical (ionic, covalent, H-bond)
- dispersive (van der Waals forces)
- diffusive (materials diffuse into each other)

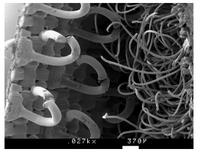
} molecular level

most frequent and general


Further types:

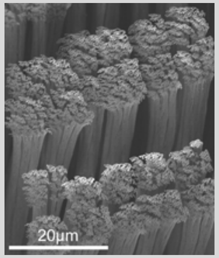
- mechanical
- electrostatic (charged surfaces)

macroscopic level



In general, the **strength of adhesion** depends on the **size and distance** of touching surfaces

„Super adhesion" of gecko



Quantitative description of adhesion

Specific interfacial energy, (σ):

$$\sigma = \frac{\Delta E}{\Delta A} \quad \left(\frac{\text{J}}{\text{m}^2} = \frac{\text{N}}{\text{m}}\right)$$

change in energy associated with the increase of surface by ΔA

increase in the interfacial area

Adhesion in dentistry:

- Surface size – acid treatment
- Viscosity
- Wetting (adhesion between solid and liquid)

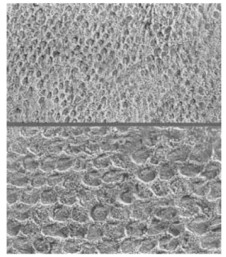
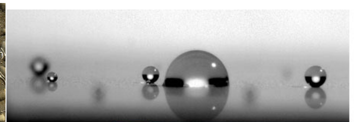
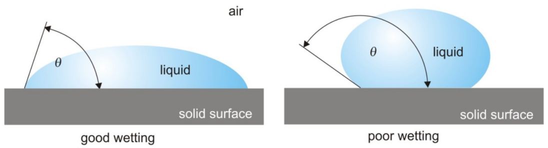


Figure 1. Morphological aspect of the surface of enamel conditioned with 36% phosphoric acid for 20 s. The formation of micropores with type I pattern of conditioning can be observed. (Original magnification: top, 750X; bottom, 1500X).

Wetting

(adhesion between solid and liquid)





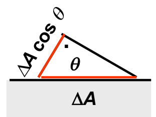
θ : contact angle

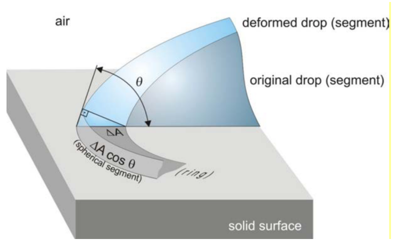
Young-equation:

$$\cos\theta = \frac{\sigma_s - \sigma_{s,l}}{\sigma_l}$$

- s : solid body (–air)
- s, f : solid body – liquid
- l : liquid (–air)

Derivation of Young's equation:

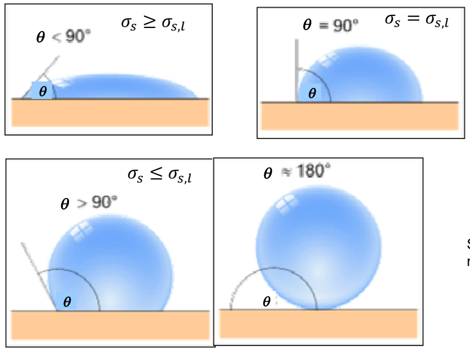




equilibrium = energy minimum ⇨ small change in shape (surface) will not cause any changes in energy

$$\Delta E = \Delta A \cdot \sigma_{s,l} - \Delta A \cdot \sigma_s + \Delta A \cdot \cos\theta \cdot \sigma_l = 0$$

$$\cos\theta = \frac{\sigma_s - \sigma_{s,l}}{\sigma_l}$$

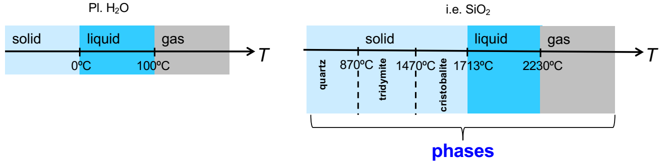


Specific surface energy of some denatl materials:

material	σ (mJ/m^2)
water	73 (25° C)
saliva	53 (37° C)
dentin	92
enamel	87
PMMA	37
paraffin	25

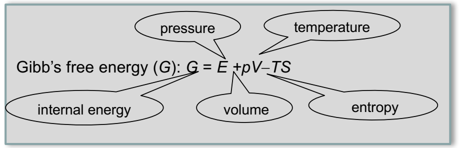
Phase

States of matter:



Phase: physically and chemically homogeneous state of a material

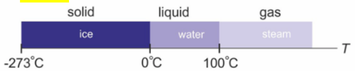
Stabile phase: The phase with the lowest Gibb's free energy at given circumstances.

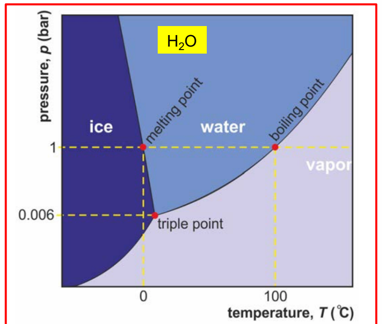


Phase diagram

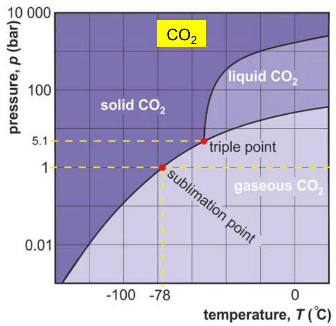
Examples:

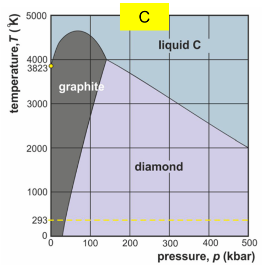
Phase diagram: plotting the stable phases as a function of different parameters (p, T, c, …)

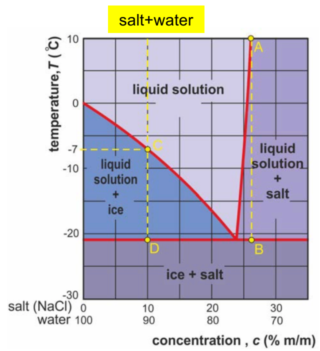


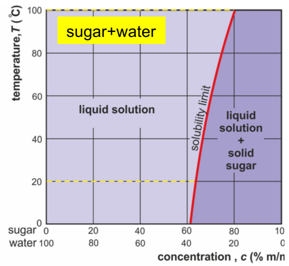


You should know only water's phase diagram for the exam.



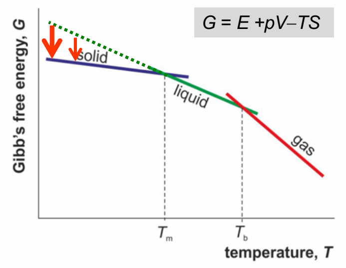


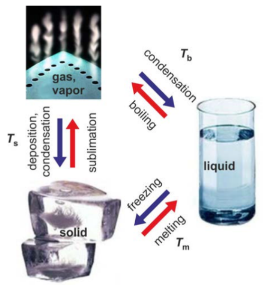




Phase transition

$G = E + pV - TS$





- Driving force: difference in free energy
  - ⇒ the smaller the T (< T_o), the greater the driving force
- Limiting factor: movement of atoms, molecules (diffusion)
  - ⇒ the smaller the T (< T_o), the weaker the movements

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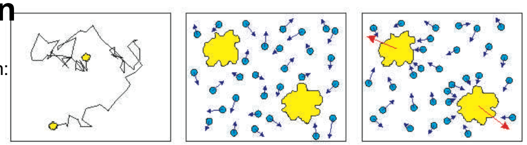
Diffusion

Brownian motion:



Diffusion: equilibration of concentration by random (thermal)

„speed" of diffusion ~ concentration difference, diffusion coefficient

Fick's law:

$$\frac{\Delta \nu}{\Delta t} = -D \cdot A \cdot \frac{\Delta c}{\Delta x}$$

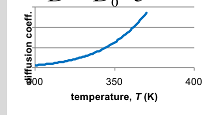
D: diffusion coefficient (m²/s)

Einstein-Stokes equation:
(for spherical particles)

$$D = \frac{kT}{6\pi\eta r}$$

$$D = D_0 \cdot e^{-\frac{\Delta\varepsilon}{kT}}$$

Diffusing molecule	medium	*D* (m²/s)
O_2	air	≈10^{-5}
	water	≈10^{-9}
	glass	≈10^{-50}
He	glass	≈10^{-18}



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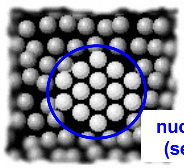
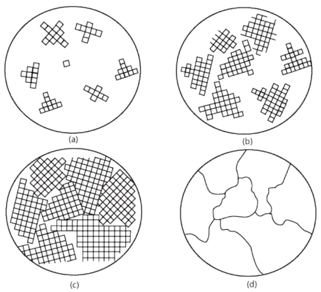
Kinetics of phase transitions (i.e. crystallization)

Overcooling! $T < T_o$

1. Nucleation



2. Growth



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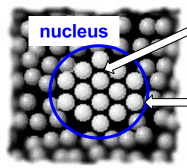
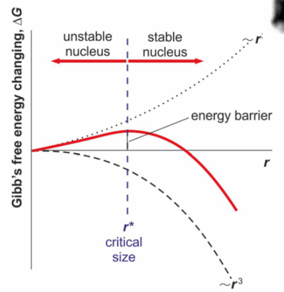
1. Nucleation

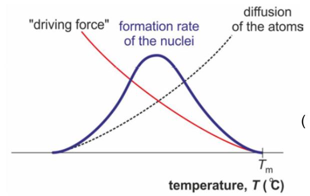


ordered structure
⇒ decrease in energy
(~ volume)
r^3

interface ⇒ increase in energy
(~ surface)
r^2

$T < T_o$:





(⇒ overcooling)

- **homogenous nucleation:** inside the same material
- **heterogeneous nucleation:** on solid surfaces (i.e. wall of container, impurity particles)

faster!

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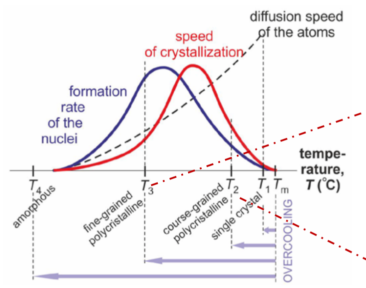
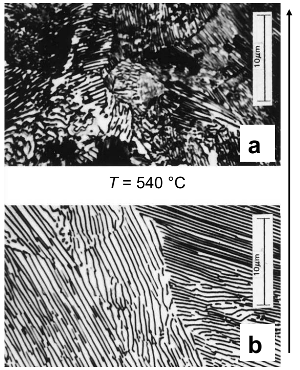
2. Crystal growth

Shape and size of grains ⇒ properties!



Example:



T = 540 °C

T = 690 °C (T_o = 727 °C)

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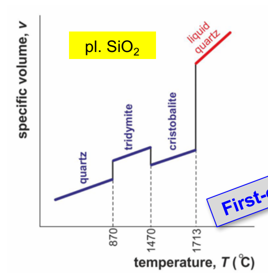
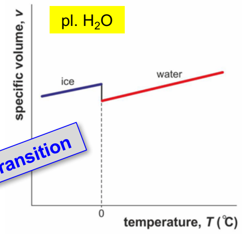
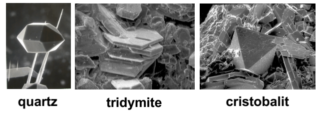
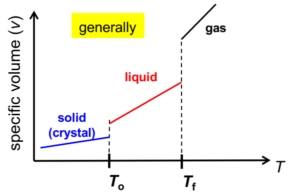
Order of phase transitions





First-order phase transition





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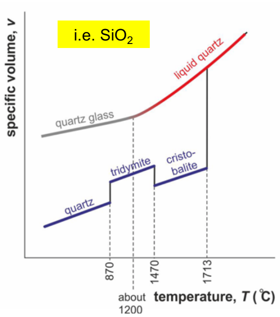
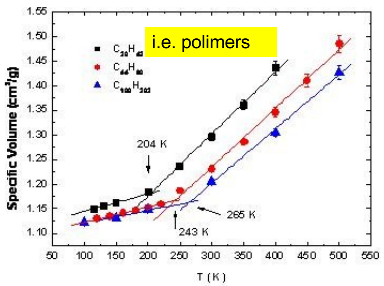
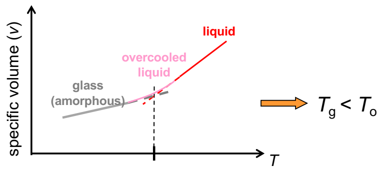




Second-order phase transition i.e. vitrification



⇒ $T_g < T_o$

T_g temp. of glass transition (vitrification)

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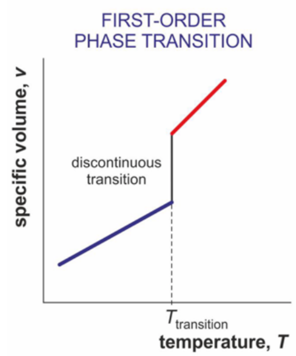
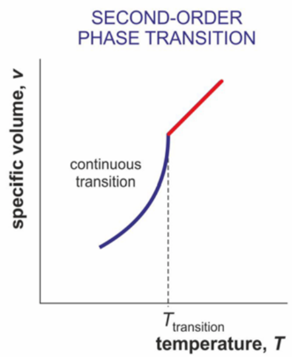
Summary:





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